

Structure and Magnetic Properties of Nanocrystalline NiFe₂O₄ Prepared via Precipitation Route

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Nanocrystalline Ni-ferrite was synthesized by modified precipitation method in which soluble starch is used as dispersing agent and Na₂CO₃ as a precipitating agent. NiSO₄·6H₂O and Fe(NO₃)₃·9H₂O were used as precursors for nickel and ferric oxide, respectively. The obtained nanocrystalline Ni-ferrite was analysed and discussed through structural, compositional and magnetic characterization. Formation of pure NiFe₂O₄ phase with average crystallite size of 21 nm has been confirmed by X-ray diffraction analysis (XRD). The determined phase composition was additionally supported by results of ⁵⁷Fe Mössbauer phase (MS) analysis and material's nanocrystalline structure by field emission scanning electron microscopy (FE-SEM). Thermomagnetic behaviour was studied up to 800 °C. The obtained room temperature magnetic hysteresis loop, recorded by means of a vibrating sample magnetometer (VSM), exhibits characteristic "S" shape of the soft magnetic material with the measured coercivity of about 10 kA/m and the specific moment up to 40 Am²/kg.

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1. Introduction

Ni-ferrites are well known as one of the most commonly used soft magnetic materials. Nevertheless, recent developments in nanotechnology have broadened their already wide range of application and opened new research areas e.g. catalysis, photoelectric and nano devices, sensors, microwave devices and magnetic pigments [1, 2]. Considering that there is a strong relation between the functional properties of Ni-ferrites and their microstructure, composition, particle and crystal sizes, the applied synthesis method plays very significant role. Hence, although synthesis of nanosized NiFe₂O₄ has been extensively studied over the years the investigation of alternative and innovative processing routes are still in progress.

2. Experimental

NiFe₂O₄ nanocrystalline powder was prepared by modified precipitation route with soluble starch (C₆H₁₀O₅)_n acting as a dispersing agent. In the first step of synthesis procedure, water solutions of analytical grade inorganic salts NiSO₄·6H₂O and Fe(NO₃)₃·9H₂O were mixed in water solution of starch. Once mixed in, they break up into fine droplets confined within the polymer chains. With addition of anhydrous Na₂CO₃ (precipitation agent) an insoluble solid phase is formed. Na₂CO₃ was added to the mixture, in a quantity calculated for possible chemical reactions. Mixture was further stirred and then dried at 80 °C until solid composite was obtained. As the starch solidifies, it entraps the formed crystals in a solid organic matrix. In the final step, the

solid composite was combusted and calcinated at 900 °C for 5 h in order to remove the organic matrix and transform reaction products to oxides which react and give NiFe₂O₄.

Microstructure of the obtained NiFe₂O₄ powder was analyzed by field emission scanning electron microscopy (FE-SEM). Phase composition and structure were analyzed by X-ray diffraction (XRD). ⁵⁷Fe Mossbauer spectroscopy (MS) was used for additional phase composition analysis. Room temperature magnetic measurements were carried out by means of a Vibrating sample magnetometer (VSM) with magnetic field strength of 1000 kA·m⁻¹. Thermomagnetic behaviour was studied up to 800 °C.

3. Results and discussion

The obtained FE-SEM image, given in Fig. 1, reveals fine-grained network structure of the prepared NiFe₂O₄ powder with particles of about 20 nm.

The results of XRD analysis for the obtained powder, point to the monophasic composition of pure NiFe₂O₄ phase (ICSD #158834). The calculated crystallite size of 21 nm is in line with the particle size observed by FE-SEM, suggesting that the prepared powder consists of single crystal particles.

The results of MS analysis (Fig. 2) support the findings of the microstructural and XRD analysis, as the presented spectrum is characteristic of Ni-ferrite material. Given that the values of Mössbauer parameters i.e. isomer shift (IS) and hyperfine magnetic field (B_{hf}) were found to be in very good agreement with the corresponding values reported for NiFe₂O₄ [3], corresponding spectral components were assigned to tetrahedral (A) and octahedral (B) Fe³⁺ atom sites. The remaining set of sextets with lowered splitting represent individual Fe posi-

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tions with diminishing iron content in the neighbourhood of both regular sites. These iron atoms are situated on crystallite surfaces and interfaces (Fig. 2), confirming the very fine crystalline structure of the material, and pointing to a variety of crystallite sizes of NiFe₂O₄ phase. The observed doublet can be associated with presence of superparamagnetic fraction of the NiFe₂O₄ phase, having size of the order of 10-13 nm or less [4].

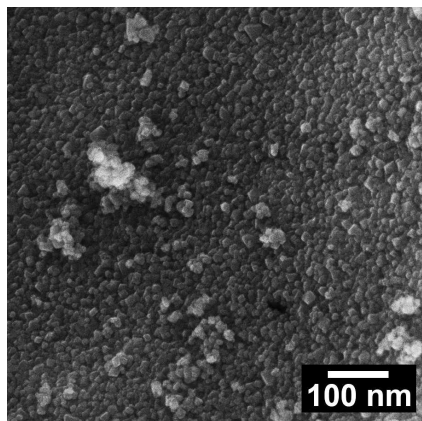


Fig. 1. Microstructure of the as-prepared NiFe₂O₄ powder.

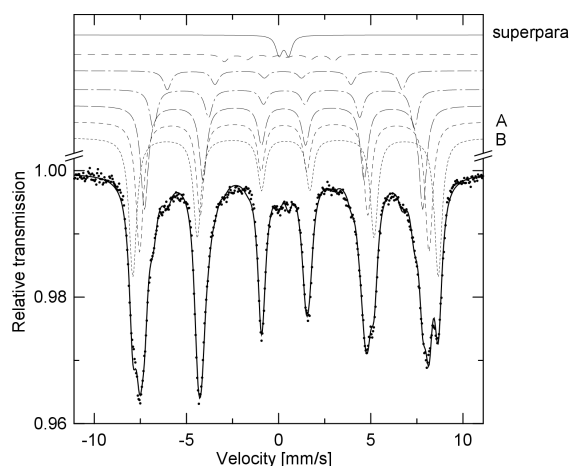


Fig. 2. Mössbauer spectrum of the obtained NiFe₂O₄ powder.

Measured values of coercivity of the order of 10 kA/m and the specific magnetic moment up to 40 Am²/kg of the prepared NiFe₂O₄ powder at room temperature are included to the range expected for this type of material.

The recorded thermomagnetic curve, presented in Fig. 3, illustrates noticeable difference in the net magnetic moment of the material before and after thermal treatment. The observed behaviour is expected and can be ascribed to fieldcooling process. In addition, it is possible that some of the remaining defects and strains were

removed by heat treatment and more ordered structure with larger crystals was obtained.

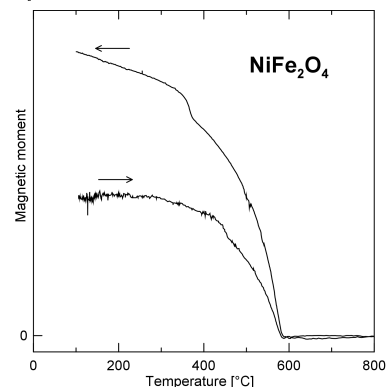


Fig. 3. Thermomagnetic curve of the NiFe₂O₄ powder.

It can be reasonably assumed that the influence of heating during thermomagnetic measurements would be reduced, if longer calcination time was applied.

4. Conclusions

The monophase nanocrystalline NiFe₂O₄ powder consisting of essentially single crystal particles with size of about 20 nm was successfully prepared by presented modified precipitation method. The results of magnetic measurements and the recorded thermomagnetic curves illustrate anticipated magnetic behaviour for this type of material. The presented synthesis method is easy and inexpensive and, unlike sonochemical [5], hydrothermal [2] or plasma route, it doesn't require any additional and expensive equipment. On the other hand, it yields material with much better uniformity of particle size than conventional coprecipitation from solution [6] or solid state reaction.

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