

ACCEPTED MANUSCRIPT

This is an early electronic version of an as-received manuscript that has been accepted for publication in the Journal of the Serbian Chemical Society but has not yet been subjected to the editing process and publishing procedure applied by the JSCS Editorial Office.

Please cite this article as: M. Zlatar, M. Gruden, *J. Serb. Chem. Soc.* (2019)

<https://doi.org/10.2298/JSC190510064Z>

This “raw” version of the manuscript is being provided to the authors and readers for their technical service. It must be stressed that the manuscript still has to be subjected to copyediting, typesetting, English grammar and syntax corrections, professional editing and authors’ review of the galley proof before it is published in its final form. Please note that during these publishing processes, many errors may emerge which could affect the final content of the manuscript and all legal disclaimers applied according to the policies of the Journal.

Calculation of the Jahn-Teller parameters with DFT

MATIJA ZLATAR and MAJA GRUDEN^{1,*}

*University of Belgrade-Institute of Chemistry, Technology and Metallurgy,
Department of Chemistry, Njegoševa 12, Belgrade, Serbia;*

¹*University of Belgrade, Faculty of Chemistry, Studentski trg 12-16, Belgrade, Serbia*

(Received 10 May; accepted 25 June 2019)

Abstract: In this review, we present Density Functional Theory (DFT) procedure to calculate the Jahn-Teller (JT) parameters in a non-empirical way, which does not depend on the system at hand. Moreover, the Intrinsic Distortion Path (IDP) model that gives further insight into the mechanism of the distortion is presented. Summarized results and their comparison to the experimentally estimated values and high-level ab initio calculations, not only proves the good ability of used approach but also gives many answers on intriguing behavior of the JT active molecules.

Keywords: vibronic coupling; density functional theory; intrinsic distortion path; distortion; transition metal complexes; organic ions and radicals

INTRODUCTION

Since its discovery more than 80 years ago, the Jahn-Teller (JT) effect¹ has drawn much attention in the scientific community, usually among scientists described as: "When a baby cries without anyone knowing why, we say - these are the teeth. Similarly, in the chemistry of transition metal compounds, when experimental phenomenon cannot be interpreted easily, it is often attributed to the Jahn-Teller effect. Moreover, the Jahn-Teller effect is at the same time a source of comfort and anxiety. Comfort because it is still there as a last resort to explain a result. Anxiety because it is challenging to reliably prove it, without remaining ambiguity".² A witty quotation, translated from French, reflects difficulties and challenges in both experimental and computational determination of the JT effect. It echoes another common misapprehension - that the JT effect is related only to the transition metal complexes, which is certainly not the case. The JT effect also occurs in highly symmetric open-shell molecules, such as cyclopentadienyl radical, benzene cation, fullerene ions, to name a few. After Nobel prize for physics in 1987,³ when it has highlighted its importance in high-Tc superconductivity, it

*Corresponding author E-mail: gmaja@chem.bg.ac.rs
<https://doi.org/10.2298/JSC190510064Z>

became more apparent that this phenomenon is not only of fundamental interest but also applicable in many areas of chemistry and physics.⁴

JT theorem states that in a non-linear molecule with degenerate ground state, structural distortion must occur, which removes the degeneracy, lowers the symmetry and makes the system more stable. In linear molecules because of symmetry constraints degenerate electronic states cannot couple with appropriate vibrations. However, these molecules can still be distorted, as described by the Renner–Teller (RT) effect.⁵ Although here we first mention the JT effect, it is worth to mention that historically it was preceded by the discovery of the RT effect.⁵ A requirement that molecule must be in a degenerate electronic state to be a subject of the JT and RT effects is excluded in the pseudo-Jahn-Teller (PJT) formalism.^{4,6,7} All these effects belong to the so-called “vibronic interactions.”^{4,8,9} Vibronic coupling is a quantum mechanical description of the influence of vibrations on the electronic structure, and vice versa. These effects are in the origin of various molecular properties such as colossal magnetoresistance,^{10,11} high-Tc superconductivity,^{3,12} unconventional superconductivity,¹³ single-molecule transport,¹⁴ ferroelectricity,^{4,15} spin-crossover,¹⁶ dynamics.⁹ They are essential in the design of single molecular magnets.^{17–19} The vibronic coupling is responsible for the observation of symmetry-forbidden transitions in spectroscopy,²⁰ conical intersections in photochemistry, and is indispensable for understanding the spectroscopy of linear molecules.^{21–26} The cause of any structural distortion of a polyatomic system is due to the JT, RT, PJT effect, or their combination.^{4,7,27} Thus, the shape of molecules, both linear²⁸ and non-linear,^{29,30} as well as solids,³¹ is explained in the framework of vibronic coupling theories. Coupling of electronic structure and nuclear movements is also essential for a proper understanding of chemical reactions.^{32–36}

Due to the coupling of electronic and vibrational motions, the Born-Oppenheimer (BO) approximation fails down at the point of electronic degeneracy. That was one of the main reasons why it has been wrongly assumed that first principle calculations cannot be applied to analyze JT or RT effects. However, the breakdown of the BO approximation brings an apparent paradox. The JT type effects emerge because of non-validity of the BO approximation, but the effects as such would not exist if the BO approximation is not used in a first place. The JT effect exists only in the realm of the BO approximation. Without the BO we do not have a picture of the molecule, and there would be no sense in discussing a distortion of a molecular structure. Standard concepts in theoretical chemistry, like force or force constants, would lose their meaning. Nonetheless, a perturbation approach restores the BO approximation. Hamiltonian is written as a Taylor series in nuclear coordinates. Coefficients in the polynomial expression of the adiabatic potential energy surface are the vibronic coupling constants, and they quantify the strength of the coupling between the electronic structure and

nuclear displacements.⁴ These coefficients are complicated integrals that have physical interpretation, like force and force constant at high symmetry configuration.⁴ Conventional computational methods can still be used if the adiabatic potential energy surface is accurately determined, which is nowadays possible with most computational methods. Many studies confirmed the excellent ability of both wave-function based methods and density functional theory (DFT) based methods to analyze the vibronic coupling.^{4,37} Typical studies do not calculate the vibronic coupling constants directly, but the calculated adiabatic potential curves along the distortion are fitted to the polynomial expression that resulted from the vibronic coupling model.^{30,38-41} Computational studies made it possible to get a more in-depth insight into the properties of the JT active compounds and understanding of many manifestations of the JT effect.

In the last decades, DFT emerges as the mainstream among computational methods, preferably because it gives a good compromise between accuracy and computational time. The theory of DFT is well documented,⁴²⁻⁴⁸ and here we will give just a brief explanation in a non-mathematical way what density functional theory is and how it can be applied to analyze the JT effect. In its theory DFT gives exact energy of the system as a method of obtaining a solution to the Schrödinger equation of a many-body system, using electron density. The Hohenberg-Kohn theorem⁴⁹ asserts that the density of any system determines all ground-state properties of the system, so the total ground state energy of a many-electron system is a functional of the density. Consequently, if we know the electron density functional, we have the exact total energy of our system. In practical computational work, as there is still no universal functional, approximation needs to be made, leading to many functionals designed for certain properties.^{43,50}

There is a controversy in the literature whether DFT can be used for the analysis of the JT effect,⁵¹⁻⁵³ as the original Hohenberg-Kohn theorems⁴⁹ were formulated only for non-degenerate ground states in the absence of a magnetic field. However, the second reformulation of this theorem gives formally proof that DFT can handle degenerate state.^{54,55} This issue is also elaborated within the sub-system ensemble DFT,^{56,57} the spin-restricted ensemble-referenced Kohn-Sham (KS) DFT,⁵⁸ the nonadiabatic generalization of DFT⁵⁹ and the sign-change in DFT.⁶⁰ Additionally, DFT naturally encompasses the mechanism of the JT effect.^{61,62} Electron density must be totally symmetric in a point group of a molecule. This is not possible if the orbitals belonging to a degenerate irreducible representation (irrep) are not equally occupied. In such a case, non-totally symmetric density will lead to the non-isotropic forces on nuclei, driving the distortion to the molecular structure with lower point-group than the initial one. Molecular structure with degenerate electronic states in the high-symmetry (HS) point group is, therefore, not a stationary point. In the case of PJT effect, change

in the distribution of electrons along a distortion is accompanied by lowering the force constant of the high-symmetry configuration.^{4,7,29,63} Admittingly, from a practical point of view, care should be taken when dealing with JT type effects with DFT.^{37,64} Pragmatic way for a detailed calculation of the JT effect is developed by Daul *et al.*,^{37,64,65} and it has been successfully applied for the analysis of many JT active molecules.^{64,65,74,75,66–73}

METHODOLOGY

Within vibronic coupling theory, Hamiltonian, \mathcal{H} , is expanded as a Taylor series around the HS molecular structure, along the normal modes, Q_i :

$$\mathcal{H} = \mathcal{H}_0 + \sum_i (\partial \mathcal{H} / \partial Q_i) Q_i + 1/2 \sum_{ij} (\partial^2 \mathcal{H} / (\partial Q_i \partial Q_j)) Q_i Q_j + \dots \quad (1)$$

where \mathcal{H}_0 is the Hamiltonian for the reference nuclear configuration, and the sum that follows it is the JT or vibronic Hamiltonian, W . Adiabatic potential energy surface, for the f -fold degenerate electronic state, takes the form

$$E_k = 1/2 \sum_i K_i Q_i^2 + \varepsilon_k \quad (2)$$

where $k = 1, 2, \dots, f$, and K_i is the force constant for the vibration Q_i and ε_k are the roots of the secular equation:

$$|W - \varepsilon I| = 0 \quad (3)$$

W is an $f \times f$ vibronic matrix, and I is a unit matrix of the same dimensions. Matrix elements of the vibronic operator are the vibronic coupling constants. For instance, the linear vibronic coupling constant has the form $F_{i,k} = \langle \Psi_k | \partial \mathcal{H} / \partial Q_i | \Psi_k \rangle$, where Ψ_k is the wavefunction belonging to the k -th component of the degenerate state. Group theory can be used to judge whether these matrix elements are different from zero. Irrep of the JT active normal modes must belong to the same representation as the symmetric direct product of the components of the degenerate electronic state, to have the linear coupling constant different from zero.

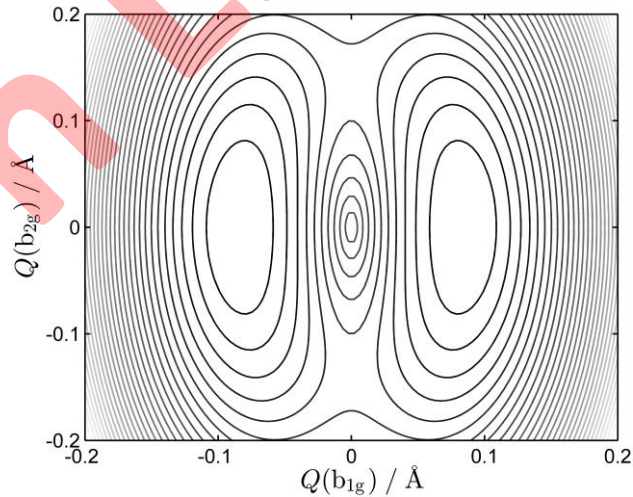


Fig. 1. Contour plot of the adiabatic potential energy surface of Sb_4^- in the space of JT active b_{1g} and b_{2g} normal modes. Minima have rectangular and transition states rhombic geometries⁷⁰

For example, in the case of the JT active Sb_4^- , electronic ground state in square planar (D_{4h}) configuration is 2E_g , and vibrations that belong to B_{1g} and B_{2g} irreps are distorting the structure to rhombus and rectangle, respectively. Both distorted structures belong to the D_{2h} point group. Adiabatic potential energy surface of Sb_4^- in the space of two JT active vibrations is depicted in Fig. 1. The parameters used to construct the surface were obtained by DFT.⁷⁰

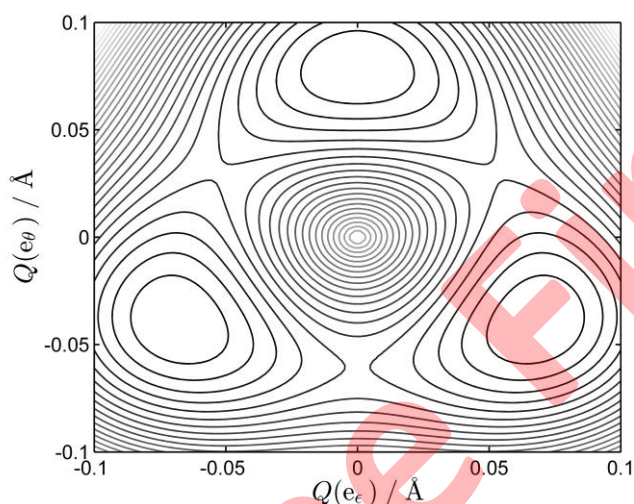


Fig. 2. Contour plot of the adiabatic potential energy surface of VCl_4 in the space of two components of the JT active e vibration. Minima have 2A_1 , and transition states 2B_1 state⁶⁴

Another example is VCl_4 tetrahedral molecule (T_d point group) with an E ground electronic state coupled with double degenerate vibrations also belonging to E irrep. Distortion lowers the symmetry to D_{2d} , and degenerate state splits into 2A_1 and 2B_1 . Adiabatic potential energy surface has the famous wrapped “Mexican-hat-like” form, with three equivalent minima and three equivalent transition states, Fig. 2.

Vibronic coupling constants determine the shape of the adiabatic potential energy surface (Fig. 1 and Fig. 2). However, from a computational perspective, it is more natural to define it with the JT stabilization energy (E_{JT}), warping barrier (Δ) and JT radius (R_{JT}). A qualitative cut through the adiabatic potential energy surface, along the JT active distortion Q_a , is given in Fig. 3. The Fig. 3 indicates how this set of the JT parameters define the adiabatic potential energy surface of the JT active molecules. The meaning of the parameters is clear - the value of E_{JT} gives energy stabilization due to the distortion, the energy difference between the minimum (min in Fig. 3) and transition state (TS in Fig. 3) is Δ , and direction and magnitude of the distortion are given with R_{JT} .

The set of JT parameters, Fig. 3, is, at least in principle, easy to determine from the first principle calculations. To get JT parameters one needs to know proper geometry, and consequently energy of the HS configuration, as well as, geometries and energies of distorted, electronically non-degenerate low symmetry (LS) structures. Experimental determination of the parameters is less straightforward. One needs to fit the experimentally obtained results to the proposed model.

The procedure for calculating the JT parameters using DFT consists of three steps:^{37,64}

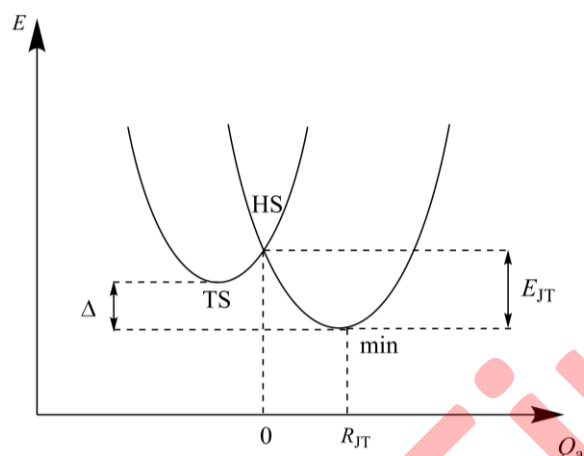


Fig. 3. Qualitative cross-section through the adiabatic potential energy surface, along the JT active distortion Q_a from the high-symmetry point (HS) toward the minimum (min) and transition state (TS); definition of the JT parameters: the JT stabilization energy (E_{JT}), the warping barrier (Δ) and the JT radius (R_{JT})

1. Geometry optimization constraining symmetry to the HS point group. This yields a geometry of the high-symmetry configuration, with an electronic structure where the electrons of degenerate orbitals are distributed equally over the components of the degenerate irreps. *E.g.*, for e^1 configuration this will mean to place 0.5 electrons into each of the two (alpha) e orbitals; for e^3 configuration, 0.5 electrons will be in each of two beta e orbitals.
2. Performing a single-point calculation constraining the HS on the nuclear geometry and the LS on the electron density. This gives the energy of a Slater determinant with an integer electron orbital occupancy. It is necessary to evaluate the energies of all possible Slater determinants with integer electron occupations. In other words, all possible modes of occupation of molecular orbitals need to be evaluated. This step is achieved by introducing an adequate occupation scheme of the molecular orbitals (MO).
3. A geometry optimization constraining the structure to the low-symmetry point group, with the proper occupancy of the KS orbitals. These calculations yield different geometries and energies that correspond to a minimum and a transition state on the adiabatic potential energy surface.

This calculation scheme is schematically drawn in the Fig. 4 for the calculation of the JT parameters in octahedral Cu(II) complexes. E_{JT} is the difference between the energies obtained in steps 2 and 3. As stated above, steps 2 and 3 are repeated for different combinations of electronic states in the LS point group. Hence there will be different E_{JT} for a minimum and transition state structures. The difference between the two E_{JT} gives the warping barrier, Δ . R_{JT} is given by the length of the distortion vector between the high symmetry and the minimum energy configuration, *i.e.*, the difference in geometries from step 1 and 3.

To avoid calculating geometry and energy at HS configuration, it is possible to use the energy of vertical (Franck-Condon) transition, E_{FC} , from the LS distorted structures.⁷⁴ This energy is easily obtained by promoting the unpaired electron from the ground state to the first

excited state for the ground state (LS) geometry. When anharmonicity is negligible, *i.e.*, in the linear and quadratic JT models, $E_{JT} = 4 E_{FC}$.

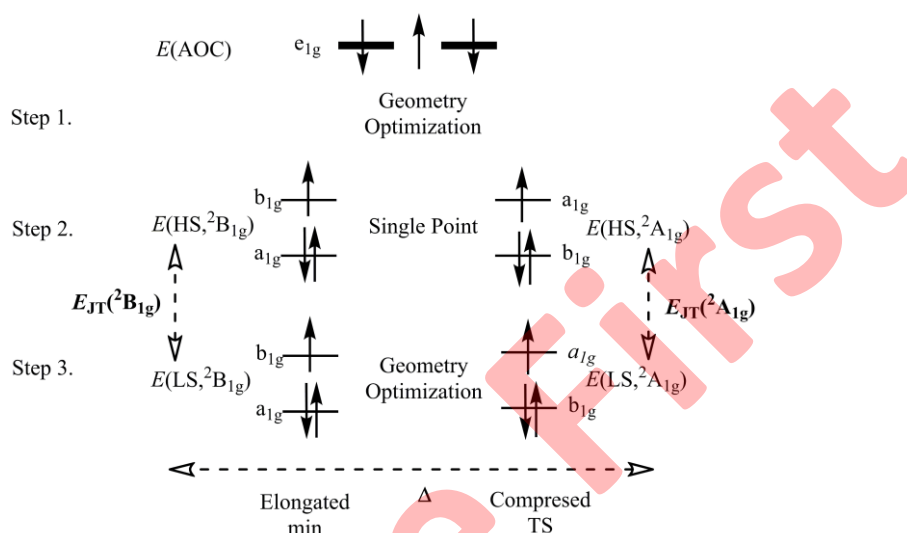


Fig. 4. DFT approach for the calculation of the JT parameters. The JT effect in octahedral Cu(II) complex is taken as an example

The first step consists of the calculation with non-integer electron configuration. Since MOs themselves have no physical meaning, using the partial occupation is just a way of obtaining electron density of the A_1 symmetry in the HS point group. Because electron density is totally symmetric, it exerts the same force on all atoms. After geometry optimization in step 1, the total force is null. Passing from this configuration with fractional occupation numbers, to the one with integer occupation, in the second and third step, leads to a change of the electronic density responsible for the non-zero force. Distortion thus occurs because a change in the electronic density requires modifying the equilibrium geometry.^{42,61} This is illustrated in Fig. 5, where the difference in electron densities between steps 2 and 1 for the case of $[\text{CuF}_6]^{4-}$ is depicted, together with the resulting force (calculated at LDA/TZP level of theory). The resulting forces lead to the tetragonal distortion and stronger equatorial bonds compared to the axial ones. Lowering of the symmetry in step 2 is necessary because this is the only way that electron density is totally symmetric with integer occupations of MOs.

Calculation of the Hessian at step 3 indicates the character of stationary points - minimum or transition state. This correlates with relative magnitudes of two E_{JT} . Interestingly, it is possible to calculate the Hessian also in step 1, and all eigenvalues should be positive. This means that geometry from step 1 is the minimum for such an electronic configuration (with fractional occupation numbers). This nuclear configuration is a minimum structure (hypothetically) in the absence of the JT effect.

The JT parameters that are obtained from the calculations, Fig. 3 and Fig. 4, relate to the vibronic coupling constants.⁴ However, exact relations will depend on the vibronic coupling model used. For example, in the case of Sb_4^+ , for each of the two JT active vibrations (b_{1g} and b_{2g}), Fig. 1, the relations are $E_{JT} = F^2/2K$ and $R_{JT} = F/K$, where F is the linear vibronic coupling

constant, and K is the force constant of the normal mode.⁷⁰ In the case of VCl_4 , Fig. 2, second-order vibronic coupling is operational, and relations are: $E_{\text{JT}}=F^2/2(K-|G|)$, $R_{\text{JT}}(\text{min})=|F|/(K-2|G|)$, $R_{\text{JT}}(\text{TS})=|F|/(K+2|G|)$, $\Delta=4E_{\text{JT}}/(K+2|G|)$, where F , K , and G are the linear vibronic coupling constant, force constant and quadratic vibronic coupling constants, respectively.⁶⁴

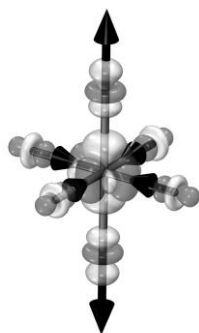


Fig. 5 Change of the electron density from the orbital configuration with fractional occupation to the one with integer occupation in O_h configuration of $[\text{CuF}_6]^{4-}$. Light/dark color represents electron density depletion/concentration (iso-value 0.002 a.u). Forces on atoms are represented as black arrows.

Calculation of the Hessian at step 3 indicates the character of stationary points - minimum or transition state. This correlates with relative magnitudes of two E_{JT} . Interestingly, it is possible to calculate the Hessian also in step 1, and all eigenvalues should be positive. This means that geometry from step 1 is the minimum for such an electronic configuration (with fractional occupation numbers). This nuclear configuration is a minimum structure (hypothetically) in the absence of the JT effect.

The JT parameters that are obtained from the calculations, Fig. 3 and Fig. 4, relate to the vibronic coupling constants.⁴ However, exact relations will depend on the vibronic coupling model used. For example, in the case of Sb_4 , for each of the two JT active vibrations (b_{1g} and b_{2g}), Fig. 1, the relations are $E_{\text{JT}}=F^2/2K$ and $R_{\text{JT}}=F/K$, where F is the linear vibronic coupling constant, and K is the force constant of the normal mode.⁷⁰ In the case of VCl_4 , Fig. 2, second-order vibronic coupling is operational, and relations are: $E_{\text{JT}}=F^2/2(K-|G|)$, $R_{\text{JT}}(\text{min})=|F|/(K-2|G|)$, $R_{\text{JT}}(\text{TS})=|F|/(K+2|G|)$, $\Delta=4E_{\text{JT}}/(K+2|G|)$, where F , K , and G are the linear vibronic coupling constant, force constant and quadratic vibronic coupling constants, respectively.⁶⁴

It is worth to mention that calculated JT parameters are not very sensitive to the choice of exchange-correlation functional.⁷³ However, the JT parameters strongly depend on geometry. Hence the proper choice of functional for the system at hand is of utmost importance. For organic molecules, this is not a difficult task, but for transition metal compounds particular caution (especially concerning spin states) is required.⁷⁶

Further insight on what is happening during distortion is based on the analysis of the R_{JT} . We already mentioned that R_{JT} describes the direction and magnitude of the distortion. In the case of triangular Na_3 cluster, it is represented by changing the angle.⁴⁰ In the other simple cases, the distortion corresponds to the movements of nuclei along one normal mode that belongs to a non-totally symmetric irrep of the HS point group of a molecule. R_{JT} is then the magnitude of the distortion along that normal mode. In simple octahedral Cu(II) complexes often the following relation is used:⁶⁵ $R_{\text{JT}}^2=\sum_i \delta d_i$, where $\delta d_i = d_i - d_0$, d_i is the distance from the Cu(II) ion to the i -th ligand, and d_0 is the average metal-ligand distance. In more complex molecules, the JT distortion is a consequence of many active modes. The evaluation of the influence of different normal modes on the JT effect is a “multimode problem,” and suitable models need to be employed to analyze the distortion.^{4,37,41,77} Two most successful models for

the analysis of the multimode problem are the “interaction mode”^{41,78,79} and “intrinsic distortion path” (IDP).^{37,67,77}

Interaction mode^{41,78,79} is a single, effective mode, that is built as a linear combination of all the JT active normal modes of the HS configuration. Coefficients in this linear combination depend on the vibronic coupling constants of each mode, and the vibronic coupling model used. Typically, interaction mode is describing the straight-line path between the HS and LS geometries, *i.e.*, its direction coincides with the direction of the R_{JT} . It is crucial to notice that the interaction mode is a reducible representation in the HS point group, while it belongs to the totally symmetric irreducible representation in the LS point group. The same is true for R_{JT} .

The essence of the IDP model^{37,67,77} is to express the distortion along a model minimal energy path starting from the DFT, or the first-principle HS geometry, and ending in the LS minimum, or the LS transition state. The reference point in IDP is the LS configuration, and the energy surface is approximated to be harmonic, *i.e.*, quadratic in the LS vibrational modes. This model allows determining the presence and contribution both to the E_{JT} and R_{JT} of all normal modes involved. All totally symmetric normal modes in the LS configuration are contributing to the distortion because IDP as a „reaction path“ must be totally symmetric.³⁶ The analytical expression of the energy allows obtaining the total force at the HS point directly. This force can be projected on either HS or LS normal modes. Knowing that the physical interpretation of the linear vibronic coupling constants is the force along an HS normal mode,⁴ IDP directly evaluates all the linear vibronic coupling constants.⁷⁷ The contributions of each mode along the path are changing, thus providing a very detailed picture of the mechanism of the distortion.

Both, interaction mode and IDP have their advantages and drawbacks, as discussed in the literature,⁷⁷ and are reducing multidimensional energy surface that is difficult to visualize to a simple cross-section.

RESULTS AND DISCUSSION

Using the procedure mentioned above, we analyzed many JT active molecules, ranging from the small metal clusters, organic radicals, organo-metallic compounds and coordination compounds (Table I). All these molecules differ in the number of atoms, nature of the chemical bonding, symmetry of the distortion, range of the JT stabilization energy, but in all cases, we obtained good agreement with experimentally evaluated values or high-level *ab initio* data.

We have already mentioned VCl_4 and Sb_4^- , Fig. 1 and Fig. 2. The modes that can contribute to the distortion are a_1 and e for VCl_4 and a_{1g} , b_{1g} , and b_{2g} for Sb_4^- . Totally symmetric modes do not change the symmetry. However, in principle, needs to be taken into consideration since their vibronic coupling constants are not zero by symmetry. Our analysis based on the IDP model indicates that totally symmetric modes do not contribute to the distortion in these molecules.^{37,70} Thus, the JT effect in VCl_4 is in essence example of the ideal, single mode JT distortion, and Sb_4^- is the simplest multimode problem. However, in Sb_4^- only one mode is responsible for the distortion to the minimum and the other one to the transition state. IDP is designed for more complicated situations⁷⁷ but is also perfectly suitable for these cases. IDP analysis justifies illustration of adiabatic

potential energy surfaces for these molecules as depicted in Fig 1. and Fig. 2. As_4^- is completely analog to Sb_4^- .⁷⁰ Vibronic coupling constants for all three molecules are calculated by IDP, by their relation to the DFT obtained JT parameters, and by fit to the potential energy curves along the JT active normal modes. Results of different approaches are consistent. DFT results for the VCl_4 ($E_{JT} = 50 \text{ cm}^{-1}$)^{37,64} are in good agreement with experiments ($E_{JT} = 30\text{-}80 \text{ cm}^{-1}$),⁸⁰⁻⁸³ and together with the values of the vibronic coupling constants, confirm the dynamical character of the JT effect.

TABLE I. Results of the DFT calculations performed to analyze the JT effect of a set of compounds, and comparison to the available reference values (ref.)

Molecule	Distortion	$E_{JT}(\text{DFT}) / \text{cm}^{-1}$	Ref. (DFT)	$E_{JT}(\text{ref.}) / \text{cm}^{-1}$	Ref.
VCl_4	$T_d \rightarrow D_{2d}$	50	37,64	30-80	80-83
Cu_3	$D_{3h} \rightarrow C_{2v}$	530	67	280-550	4
Na_3	$D_{3h} \rightarrow C_{2v}$	870	73	760-1160	40,84,85
Sb_4^-	$D_{4h} \rightarrow D_{2h}$	579	70	-	-
As_4^-	$D_{4h} \rightarrow D_{2h}$	840	70	-	-
$C_5H_5^+$	$D_{5h} \rightarrow C_{2v}$	1235	37,66	1237	86
$C_6H_6^+$	$D_{6h} \rightarrow D_{2h}$	879	66	700-1000	87-89
$C_6H_6^-$	$D_{6h} \rightarrow C_{2v}$	1187	66	-	-
$C_7H_7^+$	$D_{7h} \rightarrow C_{2v}$	853	66	1043-1374	90
C_{60}^+	$I_h \rightarrow D_{5d}$	600	77	557	91
C_{60}^-	$I_h \rightarrow C_{2h}$	300-315	92	280-600	4
$C_{20}H_{10}^+$	$C_{5v} \rightarrow C_s$	548-665	68	-	-
$C_{20}H_{10}^-$	$C_{5v} \rightarrow C_s$	404-540	68	952	93
$C_{24}H_{12}^+$	$D_{6h} \rightarrow D_{2h}$	287-401	68	484	94
$C_{24}H_{12}^-$	$D_{6h} \rightarrow D_{2h}$	290-410	68	645	94
$CoCp_2$	$D_{5h} \rightarrow C_{2v}$	870	67	1050	80,95
$MnCp_2$	$D_{5h} \rightarrow C_{2v}$	500	96	350	80,95
$[FeCp_2]^+$	$D_{5h} \rightarrow C_{2v}$	462	96	-	-
Pc^{3-}	$D_{4h} \rightarrow D_{2h}$	319-454	71	-	-
$MgPc^-$	$D_{4h} \rightarrow D_{2h}$	399-614	71	406	97
$MnPc^-$	$D_{4h} \rightarrow D_{2h}$	137-152	71	-	-
$[Cu(\text{TACN})_2]^{2+}$	$D_3 \rightarrow C_2$	1790	98	1800	99
$[Ni(\text{TACN})_2]^{3+}$ (low spin)	$D_3 \rightarrow C_2$	1504	98	1609	100
$[Co(\text{TACN})_2]^{2+}$ (low spin)	$D_3 \rightarrow C_2$	1995	98	-	-
$[Co(\text{TACN})_2]^{3+}$ (intermediate spin)	$D_3 \rightarrow C_2$	2168	98	-	-
$[Cu(\text{en})_3]^{2+}$	$D_3 \rightarrow C_2$	2200	75	1800-2200	101
$[Cu(\text{eg})_3]^{2+}$	$D_3 \rightarrow C_2$	1980	75	1620-2000	101
$[NiCl_3(\text{Hdabco})_2]^+$	$D_3 \rightarrow C_2$	253	18	-	-

In VCl_4 , for the first time, some interesting features related to the DFT scheme for the calculation of the JT parameters were observed.⁵⁸ These observations are found to be general in other cases. The first one is related to the energies of different electron distributions in step 2 (Fig. 4), which should be equal for all the possibilities. However, in practice, this is not the case due to the nature of the exchange-correlation functional involved in practical DFT calculations. The second finding is that ordering of orbitals in step 2 is *non-aufbau*. In other words, the singly occupied MO stays, after the energy minimization, above the lowest empty orbital. Expected orbital ordering is usually restored after the relaxation of the geometry (step 3). When distortion is very small, *e.g.*, in VCl_4 occupation stays even after the geometry optimization in step 3. It should be pointed out that the JT effect is a change in geometry due to the lowering of the total energy and the KS MO energies do not need to reflect that. In DFT, the total energy is not equal to the sum of the KS orbital energies. The third remark is that partial occupation of degenerate orbitals (step 1) usually leads to lower energy than the one-electron-one-orbital occupancy (step 2), because of different self-interaction errors in the two cases. In the case of VCl_4 , the energy from step 1 can be even lower than the final energies from step 3, giving a misleading result.

Cyclopentadienyl radical ($\text{C}_5\text{H}_5^\cdot$) is one of the most studied molecules from the JT perspective, and theoretical methods employed range from simple MO calculations¹⁰² to high-level *ab-initio* calculations (Table II). Studies of Mueller *et al.* who used complete active space methods ($E_{\text{JT}} = 2147 \text{ cm}^{-1}$)⁸⁶ and dispersed fluorescence spectroscopy ($E_{\text{JT}} = 1237 \text{ cm}^{-1}$),⁸⁶ as well as fitting *ab-initio* calculations to the spectra ($E_{\text{JT}} = 1463 \text{ cm}^{-1}$)⁸⁶ are considered to be benchmark results for the determination of the JT parameters in cyclopentadienyl radical. Authors also identified three dominant normal modes necessary to explain their results. Results obtained by DFT are comparable,^{37,66} or in some cases even better than other theoretical methods (Table II). With IDP model we were able to identify the three most important vibrations (C-C stretch, C-C-C bending, and C-C-H bending)⁶⁶ in agreement with the computational and experimental studies by Miller *et al.*⁸⁶

DFT results show that the energy difference between the two distorted LS structures of cyclopentadienyl radical is only around 1.5 cm^{-1} .^{37,66} Hence it is an example of a dynamic JT effect. According to DFT calculations, ${}^2\text{A}_2$ structure (en-allyl) is a transition state, while ${}^2\text{B}_1$ (dienyl) is a minimum.^{37,66} This is in agreement with CASSCF(5,5)/cc-pVTZ¹¹² and CASSCF(5,5)/6-31G*⁸⁶ calculations that report dienyl structure to be a minimum and Δ to be 4.5 and 3.6 cm^{-1} , respectively. The conversion between the ${}^2\text{A}_2$ and ${}^2\text{B}_1$ states goes around the JT cusp, Fig. 6. Due to the five-fold permutational symmetry of the cyclopentadienyl radical there needs to be five equivalent dienyl and five equivalent en-allyl structures,¹¹³ which interconvert to each other with practically no energy barrier.

TABLE II. Summary of various computational methods used to study the JT effect in cyclopentadienyl radical.

Method ^a	E_{JT} / cm^{-1}	Reference
MO	560	Liehr 1956 ¹⁰²
MO	728	Snyder 1960 ¹⁰³
MO	495	Hobey <i>et al.</i> 1960 ¹⁰⁴
HF/STO-3G	5072	Meyer <i>et al.</i> 1979 ¹⁰⁵
CI/STO-3G	2484	Borden 1979 ¹⁰⁶
HF/6-311+G*	1452	Cuhna <i>et al.</i> 1999 ¹⁰⁷
MP2/6-311+G*	3065	Cuhna <i>et al.</i> 1999 ¹⁰⁷
MP4/6-311+G*	2581	Cuhna <i>et al.</i> 1999 ¹⁰⁷
CCSD/6-311+G*	1613	Cuhna <i>et al.</i> 1999 ¹⁰⁷
CCSD(T)/6-311+G*	1613	Cuhna <i>et al.</i> 1999 ¹⁰⁷
CASSCF/cc-PVDZ	2139	Bearpark <i>et al.</i> 1999 ¹⁰⁸
CASSCF/6-31G*	2147	Miller <i>et al.</i> , 2001 ⁸⁶
CASSCF/6-31G*	1463	Miller <i>et al.</i> , 2001 ⁸⁶
CASSCF/cc-PVDZ	1665	Kiefer <i>et al.</i> 2001 ¹⁰⁹
CISD/cc-PVDZ	2553	Zilberg <i>et al.</i> 2002 ¹¹⁰
EOMIP-CCSD/DZP	1581	Stanton <i>et al.</i> 2008 ¹¹¹
DFT(LDA)/TZP	1253	Zlatar <i>et al.</i> , 2009 ³⁷
DFT (B3LYP)/TZP	1685	Zlatar <i>et al.</i> , 2013 ⁷³
IDP	1239	Zlatar <i>et al.</i> , 2010
Exp.	1237	Miller <i>et al.</i> , 2001 ⁸⁶

^aHF-Hartree-Fock; CI-Configuration Interaction; MPN-Moeller-Plesset Perturbation Theory of order N; CCSD(T)-Coupled Cluster Single, Double (Triple) excitations; CASSCF-Complete-Active-Space SCF; CISD - Single and Double excitations, single reference CI method; EOMIP-CCSD-Equation-of-motion ionization potential coupled cluster single, double excitations; LDA - Local Density Approximation; B3LYP - Becke's 3-parameters Lee-Yang-Parr hybrid functional

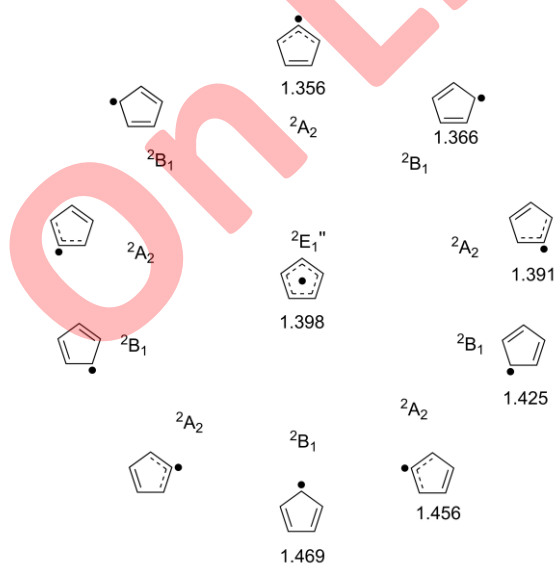


Fig. 6. Schematic representation of the pseudorotation on the lowest sheet of the adiabatic potential energy surface of the cyclopentadienyl radical. C-C bond length, Å, which is chosen to be the linear transit parameter is marked.

According to DFT calculations, 2A_2 conformations have one imaginary frequency belonging to the B_2 irrep in C_{2v} point group and lowers the symmetry to C_s . Pseudorotation of cyclopentadienyl radical goes along the path which has the C_s symmetry except at the ten configurations of C_{2v} symmetry that correspond to the stationary points on the potential energy surface. To test whether so small Δ is just due to the numerical noise in DFT calculations, the lowest sheet of the adiabatic potential energy surface has been explored using relaxed linear transit calculations. Results are shown in Fig. 7. One of the C-C bond lengths was varied from 1.469 Å, *i.e.*, single C-C bond in 2B_1 , to 1.336 Å, *i.e.*, double C-C bond in 2A_2 . Symmetry was constrained to C_s . All the remaining coordinates are allowed to be optimized. This process corresponds to the half of the pseudorotation depicted in Fig. 6.

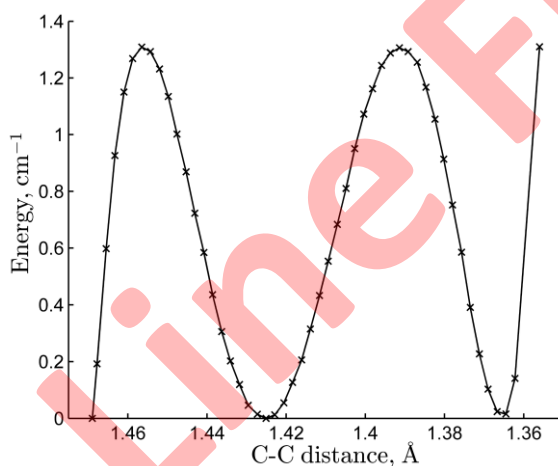


Fig. 7. Linear transit calculations used for the exploration of the lowest sheet of the potential energy surface of cyclopentadienyl radical.

In addition to the cyclopentadienyl radical, JT effect in the family of cyclic conjugated hydrocarbons was analyzed,⁶⁶ including benzene cation ($C_6H_6^+$), benzene anion ($C_6H_6^-$) and tropylium radical ($C_7H_7^{\cdot}$). All three molecules have double degenerate electronic ground state (${}^2E_{1g}$, ${}^2E_{1g}$, and $E_{2''}$, respectively) in the HS configuration (D_{6h} , D_{6h} , and D_{7h} , respectively) that is coupled with double degenerate JT active normal modes (e_{1g} , e_{1g} , and $e_{3'}$, respectively). DFT results are in excellent agreement with previous studies (Table I). IDP model showed that the C-C stretch is most important for the energy stabilization, while C-C-C bend, and C-C-H bend are contributing more to the R_{JT} . The energy component analysis along IDP revealed two distinct regions. Most of the energy stabilization is achieved by electron-nuclear attractions in the first region, where the C-C stretch is prevailing. Other forces due to the electron-electron, nuclear-nuclear,

and kinetic interactions are dominant in the second region. In the second region, little stabilization is achieved, and molecules relax toward LS stationary point via bending deformations.

In benzene cation and anion, the symmetry of the HS nuclear configuration, the symmetry of the ground electronic state, and symmetry of the JT active vibrations are all the same. However minimum on the potential energy surface is different. While in $C_6H_6^+$ minimum is D_{2h} conformation, in $C_6H_6^-$ due to the pseudo-Jahn–Teller coupling of the ground π^* electronic state with the excited σ^* state, the C_{2v} conformation is found to be the global minimum. The out-of-plane C_{2v} geometry obtained is a consequence of both JT and pseudo-JT distortion.

Corannulene ($C_{20}H_{10}$) and coronene ($C_{24}H_{12}$) and their ions are interesting molecules because besides they are building blocks for fullerene, and they also have application in electronic devices. Distortion in these molecules,⁶⁸ as in other carbohydrates always starts with the modes that are predominantly C-C stretching, and these modes contribute the most to overall distortion due to the descent in symmetry.

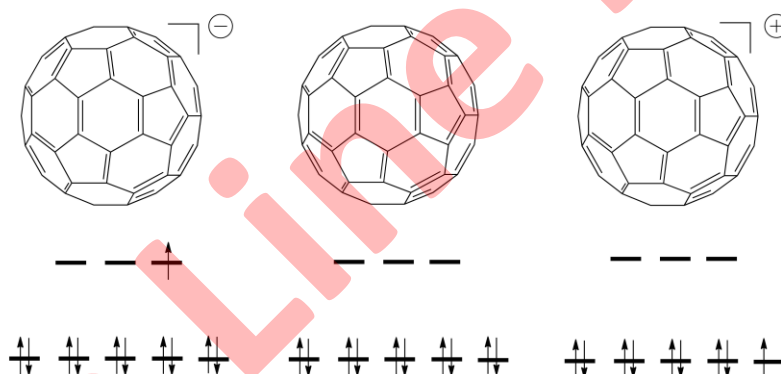


Fig. 8 Qualitative MO scheme of fullerene anion, neutral fullerene, and fullerene cation

Neutral fullerene C_{60} molecule has very high symmetry. It belongs to the I_h point group, with a high degree of degeneracies. The highest occupied molecular orbital is fivefold degenerate (h_u), and the lowest unoccupied molecular orbital is triply degenerate (t_{1u}), Fig. 8. The distortion in the JT active C_{60}^+ and C_{60}^- may lead to the structures belonging to the three different epikernel sub-groups, *i.e.*, to the structures with D_{5d} , D_{3d} or D_{2h} symmetry. The JT active modes in C_{60}^- belong to H_g irrep. Stable geometry of C_{2h} symmetry, resulting from JT distortions of double degenerate states in D_{5d} or D_{3d} , is predicted to be the most stable.⁹² In C_{60}^- the energy differences between different distorted structures are very small.⁹² According to DFT calculations, the ${}^2A_{1u}$ conformation in the D_{5d} point group is the global minimum for C_{60}^+ .⁷⁷ The vibronic coupling in C_{60}^+ is very complicated. The D_{5d} structure is a consequence of the h_g modes

described by one set of the linear vibronic coupling coefficients. D_{3d} structures are due to the distortion led by the h_g modes with a different set of vibronic coupling coefficients, but the g_g modes are active as well. Besides, totally symmetric normal modes also have non-zero vibronic coupling. Utilizing IDP model, all the vibronic coupling constants of all normal modes in C_{60}^+ were calculated⁷⁷ - two a_g , six g_g and two sets of constants for eight h_g modes. Obtained values agree with previously calculated vibronic coupling constants from the gradient of the HOMO level.⁹¹ The vibronic coupling constants of eight h_g modes and two a_g modes in C_{60}^- were also calculated⁹² by IDP model and agreed with previous reports.^{91,114} The choice of the geometry of the I_h structure was also highlighted, in particular, related to the contributions of totally symmetric normal modes. According to IDP most of the stabilization energy is achieved close to the I_h point by the hardest JT active modes. These modes have a large C–C stretch character. The relaxation of the geometry, in the final part of the distortion path, is encountered by softer modes. The situation is completely analog to the JT effect in cyclopentadienyl radical and benzene cation.

JT instable metallocenes, *e. g.*, low spin d^7 cobaltocene ($CoCp_2$, $Cp=C_5H_5$), low spin d^5 manganocene ($MnCp_2$) and low spin ferrocenyl cation, ($[FeCp_2]^+$) are typical examples of the multimode JT distortions. The HS conformation of metallocenes can be either D_{5h} if two cyclopentadienyl rings are eclipsed, or D_{5d} if two rings are staggered. According to DFT calculations eclipsed conformation is a more stable one,⁷² and descent in symmetry due to the distortion goes to C_{2v} LS structures. E_{JT} agrees with experimental results, and not differ much for three metallocenes (Table I).^{80,95} In C_{2v} minimum energy conformation, both cobaltocene and manganocene have 16 totally symmetrical normal modes which can all in principle contribute to the distortion. With IDP model their contribution to the R_{JT} , but also their contribution to the E_{JT} was determined.⁶⁷

The ground state of cobaltocene is ${}^2E_1''$ with a single electron in e_1'' MO. This MO is antibonding between metal d_{xz} (or d_{yz}) and ligand π system. Using group theory, it is easy to show that the distortion coordinate is e_2' . Out of 16 vibrations, 4 contribute more than 95% to the E_{JT} , and all four are the first order JT active, *i.e.*, e_2' , and all are located on cyclopentadienyl ring, which is not anymore planar. Although in most cases the electronic distortion due to the JT effect is mainly localized on the central metal ion, here it is delocalized over ligands. The symmetry of the electronic ground state in the HS point pushed distortion towards the perturbation of the aromaticity of Cp rings. In other words, JT induces conformational changes of the ligands, and out-of-plane ring distortion and C-H wagging minimize the antibonding interaction between single occupied d orbital of Co(II) ion and π system of Cp rings.

Manganocene with 5 electrons in d orbitals is in the ${}^2E_2'$ ground state (low spin is the most stable one) with a single hole in nonbonding e_2' orbital in D_{5h}

symmetry, with very close lying ${}^2A_1'$ state. Distortion lowers the symmetry to C_{2v} along the e_1' normal coordinates. Skeletal vibrations contribute the most to the E_{JT} and enhance small bonding interactions between d orbitals and π system of Cp rings.⁶⁷ These distortions are localized around the metal. However, IDP analysis revealed that there is one more vibrational mode which bends the two Cp rings and has an influence on the JT distortion but does not bring the dominant energetic contribution.^{67,96}

Another intriguing feature is unusual behavior of manganocene compared to its isoelectronic analog ferrocenium cation and other metallocenes. Namely, in solid state manganocene forms zig-zag polymer where the ground state is the high spin state, which was not observed in any other.¹¹⁵ Our thorough study where we applied DFT, energy decomposition analysis and IDP methods, revealed the reason behind this peculiar behavior of $MnCp_2$.⁹⁶ It has been shown that the close lying ${}^6A_1'$ state allows the conversion to the HS state present in the zigzag polymer. Moreover, the close lying excited state ${}^2A_1'$ enables the ligand deformation through the pseudo-JT coupling and triggers the polymerization. Hence, the unique behavior of manganocene has been rationalized and explained by its degenerate ground state and close lying electronic and spin states.⁹⁶

Phthalocyanine trianion (Pc^{3-}), magnesium phthalocyanine ion ($MgPc^-$) and manganese phthalocyanine ($MnPc$) are all prone to the JT distortion. Detailed analysis⁷¹ showed that coordinated metal dictated the way of breaking the symmetry: while in $MgPc^-$ the distortion is mainly localized on the ligand, in $MnPc$ central metal ion presents the trigger for the JT distortion to occur over the whole system. This difference was rationalized with IDP analysis which indicated different normal modes responsible for the JT effect in these similar systems. The linear vibronic coupling constants of all 24 JT active modes were determined by IDP model and agreed with previous work.⁷¹

Study⁹⁸ on complexes of the bis(1,4,7-triazacyclononane) (TACN) with some first-row transition metal ions with ground term subject to the JT distortion, revealed that vibronic coupling must be taken into consideration, even when the JT distortion is negligible because it can contribute to all the properties of the system. In all cases E_{JT} obtained by DFT calculations were in excellent agreement with experimentally observed ones, confirming the excellent ability of DFT approach to calculate the JT parameters even when we must deal with the complicated electronic structure of transition metal compounds. Furthermore, IDP analysis gave further insight into what is happening during the distortion and distinguished the most important normal modes out of 71 to the E_{JT} and R_{JT} . It has been shown that second coordination sphere does not influence values of the JT parameters,⁹⁸ as in the case of $[Cu(en)_3]^{2+}$ (en is ethylenediamine) and $[Cu(eg)_3]^{2+}$ (eg is ethyleneglycole).⁷⁵

Pt(PF₃)₄ is tetrahedral d⁰ molecule in ¹A₁ ground electronic state. This molecule is an important precursor for focused electron beam induced processing.¹¹⁶ In this technique, a focused beam of electrons is used to initiate the dissociation of ligands in the precursor, ideally resulting in a pure metal deposit. Several processes occur during this process, and the importance of so-called neutral dissociation is first highlighted in Pt(PF₃)₄.¹¹⁷ Neutral dissociation is a dissociation of electronically excited states initiated by incoming electrons. On first sight, this small symmetric molecule is not related to the JT effect. However, its lowest excited states are degenerate (lowest ^{1,3}T₁ and ^{1,3}T₂) and are consequently subject to the JT distortion. Distortion in these excited states is asymmetric Pt–P stretch of T₂ symmetry and this is exactly the dissociation coordinate. Therefore, the JT effect in the excited states triggers dissociation.¹¹⁷ These four excited states are directly dissociative, while other excited states are dissociative via many conical intersections.¹¹⁷

Five-coordinated Ni(II) trigonal-bipyramidal complexes (D_{3h}) in a high spin are in ³E' ground state, hence prone to the JT distortion to C_{2v} structure. It has been shown that controlling the geometry of a transition metal complex is the way to chemically control their magnetic properties, particularly that spin-orbit coupling (SOC) can be used to tune the magnetic anisotropy.^{18,19} If the SOC is higher, the magnetic anisotropy has greater value. Knowing the fact that both coordination of the ligands and the partial quenching of the SOC due to the JT distortions reduce the magnetic anisotropy from its ideal, the free ion value, we found system ([NiCl₃(Hdabco)₂]⁺, dabco is 1,4-diazabicyclo[2.2.2]-octane) where bulky dabco axial ligands prevent distortion.¹⁸ As a consequence of quenched JT, calculations revealed,¹⁸ and later it was experimentally confirmed,¹¹⁸ that this complex displays the largest magnetic anisotropy ever observed for a mononuclear Ni(II) complex.

CONCLUSION

Cast in the form of authors' review paper, this work summarizes results obtained by the special procedure in the framework of DFT, for the analysis of the JT effect in many different molecules. The theoretical background and practical computational recipe of this non-empirical and effective method are given. Treatment of the multimode JT problem with IDP approach provides the microscopic insight into the symmetry breaking process and rationalize many manifestations of the vibronic coupling. Therefore, a fast and reliable method presented here can be considered as the preferable tool for the investigation of adiabatic potential energy surface of the JT active molecules.

Acknowledgements: This work is supported by the Serbian Ministry of Education, Science and Technological Development (Grant No. 172035)

ИЗВОД
ИЗРАЧУНАВАЊЕ ЈАН-ТЕЛЕРОВИХ ПАРАМЕТАРА ПРИМЕНОМ ТЕОРИЈЕ
ФУНКЦИОНАЛА ГУСТИНЕ

МАТИЈА ЗЛАТАР И МАЈА ГРУДЕН¹

Универзитет у Београду- Институт за хемију, технологију и металургију, Центар за хемију,
Ђеишјева 12, Београд, Србија;

¹Универзитет у Београду - Хемијски факултет, Студентски брџи 12-16, Београд

У овом прегледном раду, представљена је не-емпиријска процедура за израчунавање Јан-Телерових параметара применом Теорије функционала густине, која не зависи од конкретног система који се проучава. Представљен је и модел Својственог пута дисторзије, који даје додатни увид у механизам дисторзије. Сумирани резултати и њихово поређење са експериментално процењеним вредностима, као и поређење са резултатима *ab initio* прорачуна високог нивоа, доказују тачност и велику применљивост коришћене процедуре. Такође, овде приказани рачунарски приступ даје многе одговоре на интригантне особине Јан-Телер-активних молекула.

(Примљено 10. маја; прихваћено 25. јуна 2019)

REFERENCES

1. H. A. Jahn, E. Teller, *Proc. R. Soc. A* **161** (1937) 220–235 (<https://dx.doi.org/10.1098/rspa.1937.0142>)
2. O. Kahn, *Structure électronique des éléments de transition: ions et molécules complexes*, Presses universitaires de France, 1977
3. J. G. Bednorz, K. A. Müller, *Perovskite-Type Oxides – The New Approach to High-Tc Superconductivity - Nobel Lecture*, in G. Ekspang (Ed.), *Nobel Lect. Phys. 1981–1990*, World Scientific Publishing Co, Singapore, 1993, pp. 424–457
4. I. B. Bersuker, *The Jahn-Teller Effect*, Cambridge University Press, 2006 (<https://dx.doi.org/10.1017/CBO9780511524769>)
5. R. Renner, *Z. Physik* **92** (1934) 172–193 (<https://dx.doi.org/10.1007/BF01350054>)
6. U. Öpik, M. H. L. Pryce, *Proc. R. Soc. A* **238** (1957) 425–447 (<https://dx.doi.org/10.1098/rspa.1957.0010>)
7. I. B. Bersuker, *Chem. Rev.* **113** (2013) 1351–90 (<https://dx.doi.org/10.1021/cr300279n>)
8. T. Azumi, K. Matsuzaki, *Photochem. Photobiol.* **25** (1977) 315–326 (<https://dx.doi.org/10.1021/cr300279n>)
9. H. Köppel, W. Domcke, L. S. Cederbaum, *Multimode Molecular Dynamics Beyond the Born-Oppenheimer Approximation*, in I. Prigogine, S. A. Rice (Eds.), *Adv. Chem. Physics, Vol. 57*, John Wiley & Sons, Ltd, 2007, pp. 59–246 (<https://dx.doi.org/10.1002/9780470142813.ch2>)
10. A. J. Millis, P. B. Littlewood, B. I. Shraiman, *Phys. Rev. Lett.* **74** (1995) 5144–5147 (<https://dx.doi.org/10.1103/PhysRevLett.74.5144>)
11. A. J. Millis, B. I. Shraiman, R. Mueller, *Phys. Rev. Lett.* **77** (1996) 175–178 (<https://dx.doi.org/10.1103/PhysRevLett.77.175>)
12. J. G. Bednorz, K. A. Müller, *Z. Physik B - Condensed Matter* **64** (1986) 189–193 (<https://dx.doi.org/10.1007/BF01303701>)
13. R. H. Zadik, Y. Takabayashi, G. Klupp, R. H. Colman, A. Y. Ganin, A. Potočnik, P. Jeglič, D. Arčon, P. Matus, K. Kamarás, Y. Kasahara, Y. Iwasa, A. N. Fitch, Y. Ohishi, G. Garbarino, *et al.*, *Sci. Adv.* **1** (2015) e1500059 (<https://dx.doi.org/10.1126/sciadv.1500059>)

14. M. G. Schultz, T. S. Nunner, F. von Oppen, *Phys. Rev. B* **77** (2008) 075323 (<https://dx.doi.org/10.1103/PhysRevB.77.075323>)
15. I. B. Bersuker, *Phys. Lett.* **20** (1966) 589–590 ([https://dx.doi.org/10.1016/0031-9163\(66\)91127-9](https://dx.doi.org/10.1016/0031-9163(66)91127-9))
16. P. García-Fernandez, I. B. Bersuker, *Phys. Rev. Lett.* **106** (2011) 246406 (<https://dx.doi.org/10.1103/PhysRevLett.106.246406>)
17. M. Atanasov, J. M. Zadrozny, J. R. Long, F. Neese, *Chem. Sci.* **4** (2013) 139 (<https://dx.doi.org/10.1039/c2sc21394j>)
18. M. Gruden-Pavlović, M. Perić, M. Zlatar, P. García-Fernandez, *Chem. Sci.* **5** (2014) 1453–1462 (<https://dx.doi.org/10.1039/C3SC52984C>)
19. M. Perić, A. García-Fuente, M. Zlatar, C. Daul, S. Stepanović, P. García-Fernández, M. Gruden-Pavlović, *Chem. - A Eur. J.* **21** (2015) 3716–3726 (<https://dx.doi.org/10.1002/chem.201405480>)
20. W. Domcke, H. Köppel, L. S. Cederbaum, *Mol. Phys.* **43** (1981) 851–875 (<https://dx.doi.org/10.1080/00268978100101721>)
21. M. Perić, S. D. Peyerimhoff, *J. Mol. Spectrosc.* **212** (2002) 153–161 (<https://dx.doi.org/10.1006/JMSP.2002.8534>)
22. M. Perić, S. D. Peyerimhoff, *J. Chem. Phys.* **102** (1995) 3685–3694 (<https://dx.doi.org/10.1063/1.468599>)
23. M. Mitić, R. Ranković, M. Milovanović, S. Jerosimić, M. Perić, *Chem. Phys.* **464** (2016) 55–68 (<https://dx.doi.org/10.1016/J.CHEMPHYS.2015.11.002>)
24. M. Perić, *Mol. Phys.* **105** (2007) 59–69 (<https://dx.doi.org/10.1080/00268970601129076>)
25. M. Perić, B. Ostojić, J. Radić-Perić, *Phys. Rep.* **290** (1997) 283–357 ([https://dx.doi.org/10.1016/S0370-1573\(97\)00018-5](https://dx.doi.org/10.1016/S0370-1573(97)00018-5))
26. M. Perić, S. D. Peyerimhoff, R. J. Buenker, *Mol. Phys.* **49** (1983) 379–400 (<https://dx.doi.org/10.1080/00268978300101241>)
27. I. B. Bersuker, *Spontaneous Symmetry Breaking in Matter Induced by Degeneracies and Pseudodegeneracies*, in S. A. Rice, A. R. Dinner (Eds.) *Adv. Chem. Physics*, Vol. 160, John Wiley & Sons, Ltd, 2016, pp. 159–208 (<https://dx.doi.org/10.1002/9781119165156.ch3>)
28. P. García-Fernandez, I. B. Bersuker, *Int. J. Quantum Chem.* **112** (2012) 3025–3032 (<https://dx.doi.org/10.1002/qua.24204>)
29. P. García-Fernández, J. A. Aramburu, M. Moreno, M. Zlatar, M. Gruden-Pavlovic, *J. Chem. Theory Comput.* (2014) 1824–1833 (<https://dx.doi.org/10.1021/ct4011097>)
30. Y. Liu, I. B. Bersuker, P. García-Fernandez, J. E. Boggs, *J. Phys. Chem. A* **116** (2012) 7564–7570 (<https://dx.doi.org/10.1021/jp3032836>)
31. P. García-Fernandez, J. A. Aramburu, M. Moreno, *Phys. Rev. B* **83** (2011) 174406 (<https://dx.doi.org/10.1103/PhysRevB.83.174406>)
32. R. F. W. Bader, *Mol. Phys.* **3** (1960) 137–151 (<https://dx.doi.org/10.1103/10.1080/00268976000100161>)
33. R. F. W. Bader, *Can. J. Chem.* **40** (1962) 1164–1175 (<https://dx.doi.org/10.1139/v62-178>)
34. R. F. W. Bader, A. D. Bandrauk, *J. Chem. Phys.* **49** (1968) 1666–1675 (<https://dx.doi.org/10.1063/1.1670293>)
35. R. G. Pearson, *J. Am. Chem. Soc.* **91** (1969) 4947–4955 (<https://dx.doi.org/10.1021/ja01046a001>)
36. R. G. Pearson, *Symmetry Rules for Chemical reactions*, A Willey-Interscience Publication, 1976 (<https://dx.doi.org/10.1002/bbpc.19790830620>)

37. M. Zlatar, C.-W. Schl pfer, C. Daul, *A New Method to Describe Multimode Jahn-Teller Effect Using Density Functional Theory*, in H. Koeppele, D. R. Yarkoni, H. Barentzen (Eds.), *The Jahn-Teller-Effect Fundamentals and Implications for Physics and Chemistry*, Springer Series in Chemical Physics, 97, 2009, pp. 131–165 (https://dx.doi.org/10.1007/978-3-642-03432-9_6)
38. A. R. Ilkhani, N. N. Gorinchoy, I. B. Bersuker, *Chem. Phys.* **460** (2015) 106–110 (<https://dx.doi.org/10.1016/J.CHEMPHYS.2015.07.015>)
39. A. R. Ilkhani, W. Hermoso, I. B. Bersuker, *Chem. Phys.* **460** (2015) 75–82 (<https://dx.doi.org/10.1016/J.CHEMPHYS.2015.02.017>)
40. P. Garcia-Fernandez, I. B. Bersuker, J. A. Aramburu, M. T. Barriuso, M. Moreno, *Phys. Rev. B* **71** (2005) 184110–184117 (<https://dx.doi.org/10.1103/PhysRevB.71.184117>)
41. Y. Liu, I. B. Bersuker, W. Zou, J. E. Boggs, *J. Chem. Theory Comput.* **5** (2009) 2679–2686 (<https://dx.doi.org/10.1021/ct9002515>)
42. R. G. Parr, W. Yang, *Density-Functional Theory of Atoms and Molecules*, Oxford University Press, 1989 (<https://global.oup.com/academic/product/density-functional-theory-of-atoms-and-molecules-9780195092769?cc=rs&lang=en&>)
43. C. J. Cramer, D. G. Truhlar, *Phys. Chem. Chem. Phys.* **11** (2009) 10757–10816 (<https://dx.doi.org/10.1039/b907148b>)
44. F. Neese, *Coord. Chem. Rev.* **253** (2009) 526–563 (<https://dx.doi.org/10.1016/j.ccr.2008.05.014>)
45. J. P. Perdew, A. Ruzsinszky, *Int. J. Quantum Chem.* **110** (2010) 2801–2807 (<https://dx.doi.org/10.1002/qua.22829>)
46. K. Burke, L. O. Wagner, *Int. J. Quantum Chem.* **113** (2013) 96–101 (<https://dx.doi.org/10.1002/qua.24259>)
47. A. D. Becke, *J. Chem. Phys.* **140** (2014) 18A301 (<https://dx.doi.org/10.1063/1.4869598>)
48. H. S. Yu, S. L. Li, D. G. Truhlar, *J. Chem. Phys.* **145** (2016) 130901 (<https://dx.doi.org/10.1063/1.4963168>)
49. P. Hohenberg, W. Kohn, *Phys. Rev.* **136** (1964) B864–B871 (<https://dx.doi.org/10.1103/PhysRev.136.B864>)
50. L. Goerigk, A. Hansen, C. Bauer, S. Ehrlich, A. Najibi, S. Grimme, *Phys. Chem. Chem. Phys.* **19** (2017) 32184–32215 (<https://dx.doi.org/10.1039/C7CP04913G>)
51. I. B. Bersuker, *J. Comp. Chem.* **18** (1997) 260–267 ([https://dx.doi.org/10.1002/\(SICI\)1096-987X\(19970130\)18:2<260::AID-JCC10>3.0.CO;2-M](https://dx.doi.org/10.1002/(SICI)1096-987X(19970130)18:2<260::AID-JCC10>3.0.CO;2-M))
52. I. G. G. Kaplan, *J. Mol. Struct.* **838** (2007) 39–43 (<https://dx.doi.org/10.1016/J.MOLSTRUC.2007.01.012>)
53. I. B. Bersuker, *J. Phys. Conf. Ser.* **833** (2017) 012001 (<https://dx.doi.org/10.1088/1742-6596/833/1/012001>)
54. M. Levy, *Phys. Rev. A* **26** (1982) 1200–1208 (<https://dx.doi.org/10.1103/PhysRevA.26.1200>)
55. M. Levy, *Int. J. Quantum Chem.* **110** (2010) 3140–3144 (<https://dx.doi.org/10.1002/qua.22895>)
56. A. K. Theophilou, *J. Phys. C Solid State Phys.* **12** (1979) 5419 (<https://dx.doi.org/10.1088/0022-3719/12/24/013>)
57. A. K. Theophilou, P. G. Papaconstantinou, *Phys. Rev. A* **61** (2000) 022502 (<https://dx.doi.org/10.1103/PhysRevA.61.022502>)
58. M. Filatov, *J. Chem. Theory Comput.* **9** (2013) 4526–4541 (<https://dx.doi.org/10.1021/ct400598b>)

59. R. Requist, E. K. U. Gross, *Phys. Rev. Lett.* **117** (2016) 193001 (<https://dx.doi.org/10.1103/PhysRevLett.117.193001>)
60. R. Baer, *Phys. Rev. Lett.* **104** (2010) 073001 (<https://dx.doi.org/10.1103/PhysRevLett.104.073001>)
61. J. M. García-Lastra, M. T. Barriuso, J. A. Aramburu, M. Moreno, *Chem. Phys.* **317** (2005) 103–110 (<https://dx.doi.org/10.1016/j.chemphys.2005.06.004>)
62. W. L. Clinton, B. Rice, *J. Chem. Phys.* **30** (1959) 542–546 (<https://dx.doi.org/10.1063/1.1729984>)
63. H. Nakatsuji, *J. Am. Chem. Soc.* **96** (1974) 30–37 (<https://dx.doi.org/10.1021/ja00808a005>)
64. R. Bruyndonckx, C. Daul, P. T. Manoharan, E. Deiss, R. Bruyndonckx, C. Daul, P. T. Manoharan, E. Deiss, *Inorg. Chem.* **36** (1997) 4251–4256 (<https://dx.doi.org/10.1021/ic961220+>)
65. T. K. Kundu, R. Bruyndonckx, C. Daul, P. T. Manoharan, *Inorg. Chem.* **38** (1999) 3931–3934 (<https://dx.doi.org/10.1021/ic981111q>)
66. M. Gruden-Pavlović, P. García-Fernández, L. Andjelković, C. Daul, M. Zlatar, *J. Phys. Chem. A* **115** (2011) 10801–10813 (<https://dx.doi.org/10.1021/jp206083j>)
67. M. Zlatar, M. Gruden-Pavlović, C. W. Schlöpfer, C. Daul, *J. Mol. Struct. THEOCHEM* **954** (2010) 86–93 (<https://dx.doi.org/10.1016/j.theochem.2010.04.020>)
68. L. Andjelković, M. Gruden-Pavlović, M. Zlatar, *Chem. Phys.* **460** (2015) 64–74 (<https://dx.doi.org/10.1016/j.chemphys.2015.05.007>)
69. M. Zlatar, M. Gruden-Pavlović, C.-W. Schlöpfer, C. Daul, *Chimia* **64** (2010) 161–164 (<https://dx.doi.org/10.2533/chimia.2010.161>)
70. M. Perić, L. Andjelković, M. Zlatar, C. Daul, M. Gruden-Pavlović, *Polyhedron* **80** (2014) 69–80 (<https://dx.doi.org/10.1016/j.poly.2014.02.005>)
71. L. Andjelkovic, S. Stepanovic, F. Vlahovic, M. Zlatar, M. Gruden, L. Andjelković, S. Stepanović, F. Vlahović, M. Zlatar, M. Gruden, *Phys. Chem. Chem. Phys.* **18** (2016) 29122–29130 (<https://dx.doi.org/10.1039/C6CP03859J>)
72. M. Zlatar, C.-W. Schlöpfer, E. P. Fowe, C. Daul, *Pure Appl. Chem.* **81** (2009) 1397–1411 (<https://dx.doi.org/10.1351/PAC-CON-08-06-04>)
73. L. Andjelković, M. Gruden-Pavlović, C. Daul, M. Zlatar, *Int. J. Quantum Chem.* **113** (2013) 859–864 (<https://dx.doi.org/10.1002/qua.24245>)
74. M. Atanasov, P. Comba, C. a Daul, A. Hauser, *J. Phys. Chem. A* **111** (2007) 9145–63 (<https://dx.doi.org/10.1021/jp0731912>)
75. M. Gruden-Pavlović, M. Zlatar, C.-W. Schlöpfer, C. Daul, *J. Mol. Struct. THEOCHEM* **954** (2010) 80–85 (<https://dx.doi.org/10.1016/j.theochem.2010.03.031>)
76. M. Swart, M. Gruden, *Acc. Chem. Res.* **49** (2016) 2690–2697 (<https://dx.doi.org/10.1021/acs.accounts.6b00271>)
77. H. Ramanantoanina, M. Zlatar, P. García-Fernández, C. Daul, M. Gruden-Pavlović, *Phys. Chem. Chem. Phys.* **15** (2013) 1252–1259 (<https://dx.doi.org/10.1039/c2cp43591h>)
78. V. P. Khlopın, V. Z. Polinger, I. B. Bersuker, *Theor. Chim. Acta* **48** (1978) 87–101 (<https://dx.doi.org/10.1007/BF02399020>)
79. I. B. Bersuker, V. Z. Polinger, *Vibronic interactions in Molecules and Crystals*, Springer-Verlag, Berlin, 1989 (<https://www.springer.com/gp/book/9783642834813>)
80. J. H. Ammeter, L. Zoller, J. Bachmann, P. Baltzer, E. Gamp, R. Bucher, E. Deiss, *Helv. Chim. Acta* **64** (1981) 1063–1082 (<https://dx.doi.org/10.1002/chin.198141061>)
81. F. A. Blankenship, R. L. Belford, *J. Chem. Phys.* **36** (1962) 633–639 (<https://dx.doi.org/10.1063/1.1732585>)

82. R. B. Johannesen, G. A. Candela, T. Tsang, *J. Chem. Phys.* **48** (1968) 5544–5549 (<https://doi.org/10.1063/1.1668254>)
83. Y. Morino, H. Uehara, *J. Chem. Phys.* **45** (1966) 4543–4550 (<https://dx.doi.org/10.1063/1.1727535>)
84. H. von Busch, V. Dev, H.-A. Eckel, S. Kasahara, J. Wang, W. Demtröder, P. Sebald, W. Meyer, *Phys. Rev. Lett.* **81** (1998) 4584–4587 (<https://dx.doi.org/10.1103/PhysRevLett.81.4584>)
85. J. Gaus, K. Kobe, V. Bonacic-Koutecky, H. Kuehling, J. Manz, B. Reischl, S. Rutz, E. Schreiber, L. Woeste, *J. Phys. Chem.* **97** (1993) 12509–12515 (<https://dx.doi.org/10.1021/j100150a011>)
86. B. E. Applegate, T. A. Miller, *J. Chem. Phys.* **114** (2001) 4855 (<https://dx.doi.org/10.1063/1.1348275>)
87. B. E. Applegate, T. A. Miller, *J. Chem. Phys.* **117** (2002) 10654 (<https://dx.doi.org/10.1063/1.1520531>)
88. V. Perebeinos, P. B. Allen, M. Pederson, *Phys. Rev. A* **72** (2005) 12501–12506 (<https://dx.doi.org/10.1103/PhysRevA.72.012501>)
89. K. Tokunaga, T. Sato, K. Tanaka, *J. Chem. Phys.* **124** (2006) 154303 (<https://dx.doi.org/10.1063/1.2184317>)
90. I. Sioutis, V. L. Stakhursky, G. Tarczay, T. A. Miller, *J. Chem. Phys.* **128** (2008) 084311 (<https://dx.doi.org/10.1063/1.2829471>)
91. N. Manini, A. Dal Corso, M. Fabrizio, E. Tosatti, *Philos. Mag. B* **81** (2001) 793–812 (<https://dx.doi.org/10.1080/13642810110062663>)
92. H. Ramanantoanina, M. Gruden-Pavlovic, M. Zlatar, C. Daul, *Int. J. Quantum Chem.* **113** (2013) 802–807 (<https://dx.doi.org/10.1002/qua.24080>)
93. Tokio Yamabe, Kazuyuki Yahara, and Takashi Kato, K. Yoshizawa, *J. Phys. Chem. A* **104** (2000) 589–595 (<https://dx.doi.org/10.1021/jp992496g>)
94. T. Kato, K. Yoshizawa, *J. Chem. Phys.* **110** (1999) 249–255 (<https://dx.doi.org/10.1063/1.478100>)
95. R. Bucher, ESR-Untersuchungen an Jahn-Teller-Aktiven Sandwichkomplexen, ETH Zuerich, 1977 (<https://dx.doi.org/10.3929/ethz-a-000150322>)
96. S. Stepanović, M. Zlatar, M. Swart, M. Gruden, *J. Chem. Inf. Model.* (2019) (<https://doi.org/10.1021/acs.jcim.8b00870>)
97. J. Tóbiš, E. Tosatti, *J. Mol. Struct.* **838** (2007) 112–115 (<https://doi.org/10.1016/J.MOLSTRUC.2006.12.051>)
98. M. Zlatar, M. Gruden-Pavlovic, M. Guell, M. Swart, M. Gruden-Pavlović, M. Güell, M. Swart, *Phys. Chem. Chem. Phys.* **15** (2013) 6631–6639 (<https://dx.doi.org/10.1039/C2CP43735J>)
99. P. Chaudhuri, K. Oder, K. Wieghardt, J. Weiss, J. Reedijk, W. Hinrichs, J. Wood, A. Ozarowski, H. Stratemaier, D. Reinen, *Inorg. Chem.* **25** (1986) 2951–2958 (<https://dx.doi.org/10.1021/ic00237a007>)
100. K. Wieghardt, W. Walz, B. Nuber, J. Weiss, A. Ozarowski, H. Stratemeier, D. Reinen, *Inorg. Chem.* **25** (1986) 1650–1654 (<https://dx.doi.org/10.1021/ic00230a025>)
101. E. Gamp, ESR-Untersuchungen über den Jahn-Teller-Effekt in oktaedrischen Kupfer(II)-Komplexen mit trigonalen dreizahnigen Liganden, ETH Zurich, 1980 (<https://dx.doi.org/10.3929/ethz-a-000215808>)
102. A. D. Liehr, *Z. Für Phys. Chemie* **9** (1956) 338–354 (<https://dx.doi.org/10.1524/zpch.1956.9.5.6.338>)
103. L. C. Snyder, *J. Chem. Phys.* **33** (1960) 619–621 (<https://dx.doi.org/10.1063/1.1731211>)

104. W. D. Hobey, A. D. McLachlan, *J. Chem. Phys.* **33** (1960) 1703–1965 (<https://dx.doi.org/10.1063/1.1731485>)
105. R. Meyer, F. Graf, T. Ha, H. H. Gunthard, *Chem. Phys. Lett.* **66** (1979) 65–71 ([https://dx.doi.org/10.1016/0009-2614\(79\)80370-X](https://dx.doi.org/10.1016/0009-2614(79)80370-X))
106. W. T. Borden, E. R. Davidson, *J. Am. Chem. Soc.* **101** (1979) 3771–3775 (<https://dx.doi.org/10.1021/ja00508a012>)
107. C. Cunha, S. Canuto, *J. Molec. Struct THEOCHEM* **464** (1999) 73–77 ([https://dx.doi.org/10.1016/S0166-1280\(98\)00536-3](https://dx.doi.org/10.1016/S0166-1280(98)00536-3))
108. M. J. Bearpark, M. A. Robb, N. Yamamoto, *Spectrochim. Acta A* **55** (1999) 639–646 ([https://dx.doi.org/10.1016/S1386-1425\(98\)00267-4](https://dx.doi.org/10.1016/S1386-1425(98)00267-4))
109. J. H. Kiefer, R. S. Tranter, H. Wang, A. F. Wagner, *Int. J. Chem. Kin.* **33** (2001) 834–845 (<https://dx.doi.org/10.1002/kin.10006>)
110. S. Zilberg, Y. Haas, *J. Am. Chem. Soc.* **124** (2002) 10683–10691 (<https://dx.doi.org/10.1021/ja026304y>)
111. T. Ichino, S. W. Wren, K. M. Vogelhuber, A. J. Gianola, W. C. Lineberger, J. F. Stanton, *J. Chem. Phys.* **129** (2008) 084310 (<https://dx.doi.org/10.1063/1.2973631>)
112. D. Leicht, M. Kaufmann, G. Schwaab, M. Havenith, *J. Chem. Phys.* **145** (2016) 074304 (<https://dx.doi.org/10.1063/1.4960781>)
113. A. D. Liehr, *J. Phys. Chem.* **67** (1963) 389–471 (<https://dx.doi.org/10.1021/j100796a043>)
114. N. Iwahara, T. Sato, K. Tanaka, L. F. Chibotaru, *Phys. Rev. B* **82** (2010) 245409 (<https://dx.doi.org/10.1103/PhysRevB.82.245409>)
115. J. H. Ammeter, R. Bucher, N. Oswald, *J. Am. Chem. Soc.* **96** (1974) 7833–7835 (<https://dx.doi.org/10.1021/ja00832a049>)
116. I. Utke, A. Götzhäuser, *Angew. Chemie Int. Ed.* **49** (2010) 9328–9330 (<https://dx.doi.org/10.1002/anie.201002677>)
117. M. Zlatar, M. Allan, J. Fedor, *J. Phys. Chem. C* **120** (2016) 10667–10674 (<https://dx.doi.org/10.1021/acs.jpcc.6b02660>)
118. K. E. R. Marriott, L. Bhaskaran, C. Wilson, M. Medarde, S. T. Ochsenbein, S. Hill, M. Murrie, *Chem. Sci.* **6** (2015) 6823–6828 (<https://dx.doi.org/10.1039/C5SC02854J>)